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a. If you use 2.3 moles of Al, how many moles of  $\text{Al}_2\text{O}_3$  can you make?  
b. If you want 3.9 moles of

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At 20 3, how many moles of O 2 are needed? 2. In the presence of sulfuric acid, metallic iron forms iron(III) sulfate:  $2\text{Fe} + 3\text{H}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2$  ...

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of moles of P<sub>4</sub>O<sub>10</sub> to moles of P  
which came from ...

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Answer:  $4.93 \times 10^{-5}$  L or 49.3  $\mu$ L In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent:  $[\text{Au}(\text{CN})_2]^-$ ,  $\text{LaCl}_3$ , ethanol, and para -nitrophenol. When the limiting reactant is not apparent, we can determine which reactant is limiting by comparing the molar

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instantaneous rate is the rate of a  
reaction at any particular point in time,  
a period of time that is so short that



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stoichiometry? 1. true 2. true 3. false

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4. true 5. true 6. 2, 2, 64.10 7. 3, 3,  
96.00 8. 2, 2, 88.02 9. 4, 4, 72.08 10.  
methanol and oxygen gas 11. carbon  
dioxide and water 12. 160.10 g 13.  
160.10 g 14. They are equal. 15. A  
mole ratio is a ratio between the  
numbers of moles of any two  
substances in a ...

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