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NAWAZ KHAN Chapter 11 Chemical Reactions D

Section 11.1 Assessment. Describe the steps in writing a balanced chemical equation. Write the skeleton equation for the following reactions: Heating solid copper(II)sulfide in the presence of oxygen gas produces pure copper and sulfur dioxide gas. Iron metal and chlorine gas react to form solid iron(III)chloride.  $\text{CuS (s) + O}_2\text{(g) \rightarrow Cu (s) + SO}_2\text{(g)}$  D

Chapter 11: Chemical Reactions

Chapter 11: Chemical Reactions Study Guide □□chemical equation A representation of a chemical reaction with reactants on the left, products on the right, and an arrow separating the two skeleton

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www.ck12.org Chapter 11. Chemical Reactions 11.1 Chemical Equations Lesson Objectives □ Describe chemical reactions using word equations. □ Know the correct symbols to use in order to write skeleton equations for chemical reactions. □ Use coefficients to balance chemical equations so that the law of conservation of mass is followed.

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Section 11.1 Describing Chemical Reactions. The reactants are written on the left and the products on the right. The arrow that separates them is called yield. Reactants Products. Symbols in Equations. Symbol Meaning " yields D reversible reaction (s) solid (l) liquid (g) gas (aq) aqueous " catalyst " heat. Pt .

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To describe a chemical reaction. To calculate the quantities of compounds produced or consumed in a chemical reaction. As shown in Figure 11.3.1, applying a small amount of heat to a pile of orange ammonium dichromate powder results in a vigorous reaction known as the ammonium dichromate volcano.

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balanced equation. A \_\_\_\_\_ is an equation where each side of the equation has the same number of atoms of each element, and mass is conserved. word equation. Iron + oxygen  $\rightarrow$  iron (III) oxide is an example of a(n) \_\_\_\_\_. chemical equation.  $Fe + O_2 \rightarrow Fe_2O_3$  is an example of a(n) \_\_\_\_\_.

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## Chapter 11 Chemical Reactions Vocabulary Review Answer Key

Chemistry chapter 11 Chemical reactions answer key. coefficient. a whole number that appears before a formula in an equation. spectator ion. a particle not directly involved in a chemical reaction. combustion reaction. a reaction in which oxygen reacts with another substance, often producing light or heat. reactant.

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Stoichiometry Problems. When we carry out a reaction in either an industrial setting or a laboratory, it is easier to work with masses of substances than with the numbers of molecules or moles. We will first present the method used in most other books for converting from the mass of any reactant or product to the mass of any other reactant or product using a balanced chemical equation outlined ...

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