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Stoichiometry Name

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Ammonia is made  
industrially by reacting  
nitrogen and hydrogen

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under pressure, at high temperature, and in the presence of a catalyst.

The equation is  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$ . If 4.0 mol of  $\text{H}_2$  react, how many moles of  $\text{NH}_3$  will be produced?

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OF AN ELEMENT TO

MASS Sample Problem

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of steps 1–3 to convert from the mass of an element to the number of atoms. In order to calculate the mass from a given number of atoms, these steps will be reversed.

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chemists count

quantities of elements  
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of moles and that they  
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presence of a catalyst.

The equation is  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$ . If

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