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The plot of $V = V(P, T)$ for a gas is shown in Fig. 1.1. Determine the expressions of the two second derivatives of the volume of this plot. (note: the principle curvatures of the surface are proportional to these second derivatives).

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Work is found from the first law as $w = q - \Delta U$; thus $q = \Delta H$; $w = \Delta H - P\Delta V$; 4. Isothermal Process Because U is a function only of T for an ideal gas, $\Delta U = \Delta H = 0$ for an isothermal process. These results also follow from the general results by using $DT = D(PV) = 0$ for an isothermal process.

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The isothermal expansion is conducted at 300 K. 3.4 $\Delta H = 42750$ J, $\Delta S = 59.7$ J/K 3.5 The final temperature is 323.32 K, which is greater than 323 K because the heat capacity increases with increasing temperature.

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