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Chemistry Notes for class 12 Chapter 2 Solutions Solution is a homogeneous mixture of two or more substances in same or different physical phases. The substances forming the solution are called components of the solution. On the basis of number of components a solution of two components is called binary solution. Solute and Solvent

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So, go ahead and check the Important Notes for Class 12 Chemistry Solutions. Solution is a homogeneous mixture of two or more substances in same or different physical phases. The substances forming the solution are called components of the solution. On the basis of number of components a solution of two components is called binary solution.

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1. A solution is a homogeneous mixture of two or 9. more chemically non-reacting substances. The components of a solution generally cannot be separated by filtration, settling or centrifuging. 2. A solution may be classified as solid, liquid or a gaseous solution. 3.

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CBSE Class 12th Chemistry Notes (Based on NCERT Textbook ...

Solution Chemistry Notes. 2. Solutions • Basic definitions - Solution: homogenous mixture - Solute: substance that is dissolved - Solvent: substance that contains the solute in a solution • Process of Dissolving: - Molecules of the solvent "pull apart" the solute molecules, and distributes them within the solvent • How to tell the solute from the solvent - The substance that is in the greater amount is the solvent - The substance that keeps its phase is the solvent.

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CBSE Class 12 Chemistry Notes: Solutions - Basics. by Anand Meena. April 22, 2019. in 12th Class, Class Notes. Reading Time: 2min read 0. 4. SHARES. 0. VIEWS. Share on Facebook Share on Twitter. Solution is a homogeneous mixture of two or more chemically non reacting substances whose composition can be varied within certain limits. A solution ...

CBSE Class 12 Chemistry Notes: Solutions - Basics ...

Solutions Class 12 Notes Chapter 2. A solution comprises a solute and a solvent. It is defined as a homogeneous mixture of two or more than two substances. They can be classified into three types: Solid solutions. Gaseous solutions. Liquid solutions. Molarity, mole fraction, percentages, and molality are the terms used to express the concentration of a solution.

CBSE Class 12 Chemistry Notes Chapter 2 Solutions

Free PDF download of Class 12 Chemistry revision notes & short key-notes for Chapter 2 - Solutions to score high marks in exams, prepared by expert Chemistry teachers from latest edition of CBSE (NCERT) books.

Class 12 Chemistry Revision Notes for Chapter 2 - Solutions

Revision Notes on Solution: A solution is a homogeneous mixture of two (or more) substances, the composition of which may vary between certain limits. A solution consisting of two components is called binary solution. The component which is present in large quantity is called solvent and the component which is small in quantity is called solute. If both components are in same physical state.

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CBSE Class 12 Chemistry Syllabus 2016 - 2017. Different Methods of Expressing Concentration of Solutions. The composition of solution is defined in terms of concentration.

CBSE Class 12th Chemistry Notes: Solutions (Part - I)

NCERT Solutions for Class 12 Chemistry (Chapter-Wise) NCERT Solutions for Class 12 Chemistry (all chapters) are provided chapter-wise on this page. Students can access solutions for each individual Class 12 chemistry chapter by following the links tabulated below.

NCERT Solutions for Class 12 Chemistry (All Chapters) with PDF

NCERT Notes For Class 12 Chemistry Chapter 2: Solutions. Solution is a homogeneous mixture of two or more substances in same or different physical phases. The substances forming the solution are called components of the solution. On the basis of number of components a solution of two components is called binary solution. Solute and Solvent

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The NCERT books are one of the most important resources for every class 12 student. The book 'Errorless NCERT Solutions with 100% Reasoning Class 12 Chemistry' is exclusively written to provide best quality solutions for NCERT Chemistry class 12. • The Unique Selling Point of this book lies in its quality of solutions which provides 100% Reasoning (which is missing in Most of the Books) and are Errorless. • A lot of solution provide Notes immediately after the Solutions which provides Important Tips, Shortcuts, Alternative Methods, Points to Remember etc.. • This book provides Quick Revision of the concepts involved along with Important formulas and definitions, in each chapter, which would act as a refresher. • This is followed by the detailed solutions (Question-by-Question) of all the questions/ exercises provided in the NCERT book. • The solutions have been designed in such a manner (Step-by-Step) that it would bring 100% Concept Clarity for the student. • The solutions are Complete (each and every question is solved), Inflow (exactly on the flow of questions in the NCERT book) and Errorless.

A. Surface Chemistry 1.To prepare colloidal solution (sol) of starch, 2. To prepare a colloidal solution of egg albumin 3.To prepare colloidal solution of gum, 4. To prepare colloidal solution of aluminium hydroxide [Al(OH)3], 5.To prepare colloidal solution of ferric hydroxide [Fe(OH)3], 6.To prepare colloidal solution of arsenious sulphide [As2S3], 7. To purify a freshly prepared sol by dialysis, 8. To compare the effectiveness of different common oils (Castor oil, cotton seed oil, coconut oil, kerosene oil, mustard oil) in forming emulsions. Viva-Voce B. Chemical Kinetics 1. To study the effect of concentration on the rate of reaction between sodium thiosulphate and hydrochloric acid, 2. To study the effect of temperature on the rate of reaction between sodium thiosulphate and hydrochloric acid, 3. To study the rate of reaction of iodide ions with hydrogen peroxide at different concentrations of iodide ions, 4. To study the rate of reaction between potassium iodate (KIO3) and sodium sulphite (Na2SO3) using starch solution as indicator Viva-Voce C. Thermochemistry 1.Determine the enthalpy of dis solution of copper sulphate (CuSO4.5H2O) in water at Room temperature, 2. To determine the enthalpy of neutralization of the reaction between HCl and NaOH, 3. To determine enthalpy change during the interaction between acetone and chloroform Viva-Voce D. Electrochemistry 1.To study the variation of cell potential in Zn|Zn2+||Cu2+|Cu, with change in concentration of electrolytes (CuSO4 or ZnSO4) at room temperature Viva-Voce E.Chromatography 1.To separate the coloured components (pigment) present in the given extract of leaves and flowers by ascending paper chromatography and find their Rf values, 2. To separate the coloured components present in the mixture of red and blue inks by ascending paper chromatography and find their Rf values, 3.To separate Co2+ and Ni2+ ions present in the given mixture by using ascending paper chromatography and determine their Rf values Viva-Voce F. Preparation of Inorganic Compounds 1.Preparation of double salt of ferrous ammonium sulphate (Mohr's salt) from ferrous sulphate and ammonium sulphate, 2. To prepare a pure sample of potash alum (fitkari), 3. Preparation of crystals of potassium ferric oxalate or potassium trioxalato ferrate (III) Viva-Voce G. Preparation of Organic Compounds 1. Preparation of iodoform from ethyl alcohol or acetone, 2. Preparation of acetanilide in laboratory, 3. Preparation of b-Naphthol aniline dye, 4. To prepare a pure sample of dibenzalacetone, 5. To prepare a pure sample of p-nitro acetanilide Viva-Voce H. Tests for the Functional Groups Present in Organic Compounds Viva-Voce I. Study of Carbohydrates, Fats and Proteins 1.To study simple reactions of carbohydrate, 2. To study simple reactions of fats, 3. To study simple reactions of proteins, 4. To investigate presence of carbohydrates, fats and proteins in food stuffs Viva-Voce J. Volumetric Analysis 1. To prepare 250 ml of M/10 solution of oxalic acid, 2.To prepare 250 ml of M/10 solution of ferrous ammonium sulphate, 3. Prepare M/20 solution of oxalic acid, with its help find out the molarity and strength of the given solution of potassium permanganate, 4.Prepare M/20 solution of Mohr's salt, using this solution determine the molarity and strength of potassium permanganate solution Viva-Voce K. Qualitative Analysis Viva-Voce INVESTIGATORY PROJECTS 1.To study the presence of oxalate ions in guava fruit at different stages of ripening. 2. To study the quantity of casein present in different samples of milk. 3.Preparation of soyabean milk and its comparison with natural milk with respect to curd formation, effect of temperature etc.4.To study the effect of potassium bisulphite as food preservative at various concentrations. 5. To study the digestion of starch by salivary amylase and the effect of pH and temperature on it. 6. To study and compare the rate of fermentation of the following materials-wheat flour, gram flour, potato juice and carrot juice. 7.To extract essential oils present in saunf (aniseed), ajwain (corum), illaichi (cardomom).8. To detect the presence of adulteration in fat, oil and butter, 9.To investigate the presence of NO2- in brinjal.

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